







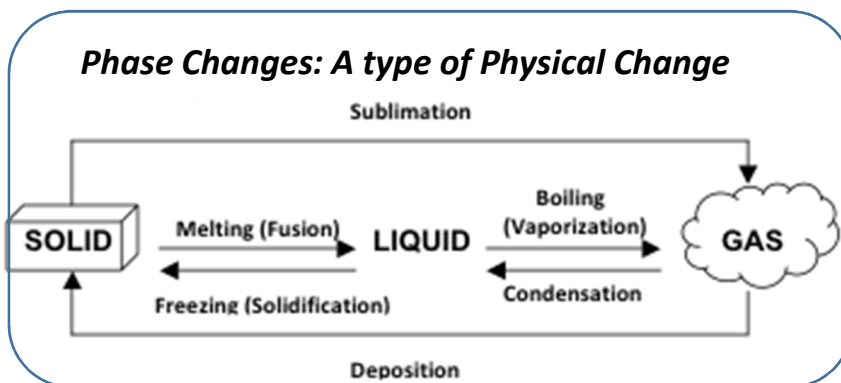


Changes we observe about matter can be categorized as either a physical or chemical change:

<p>Physical Change: a change in which the chemical composition _____ _____!</p>	<p>Chemical Change: a change in which the chemical composition _____ _____!</p>
<p>Signs of a Physical Change:</p>	<p>Signs of a Chemical Change:</p>
	
	
	
	
	
<p>***</p>	<p>***</p>



Practice: Physical or Chemical Change?

Melting butter _____
Cutting up food _____
Digesting food _____
Grating cheese _____

Fermenting milk to make cheese

Tearing a piece of Aluminum foil

Both physical and chemical changes can involve **changes in energy**

1. Consider an ice cube melting on a countertop

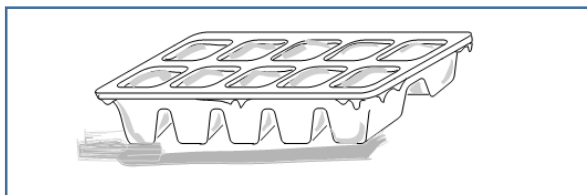
- a. Draw arrows showing which direction heat is transferred. Is heat being transferred:
- into the ice cube, or
 - out of the ice cube?



- b. When the ice cube melts, is it releasing heat or taking in (absorbing) heat from the surrounding air?

2. Consider water freezing in an ice cube tray in the freezer:

- a. Draw arrows showing which direction heat is being transferred when the ice cube tray is placed in the freezer. Is heat being transferred into or out of the water as it freezes into ice?



- b. When water freezes, is it releasing heat or taking in (absorbing) heat from the environment?

- c. In order to cool things in a refrigerator, heat must be transferred away from the food. Where do you think this heat ends up going?

- **endothermic**=when energy/heat is being _____ in a process
 - A reaction container will feel _____ due to heat being transferred into (heat is absorbed by) the reaction container
- **exothermic**=when energy/heat is being _____ in a process
 - A reaction container will feel _____ due to heat being transferred out of (heat is released by) the reaction container

The **rate of a physical or chemical change** depends on the following factors:

- **As the Temperature** _____
- **For solids: Surface Area of Reactant (Relative Particle Size)**
 - As the reactant particle size _____
- a.k.a:
 - As the surface area of the reactant _____
- **As the Concentration of the Reactants** _____
- **Presence of a Catalyst**
 - **Catalyst** = _____
 - Digestive enzymes are catalysts!

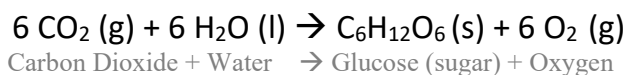
- 1) Given the balanced equation representing a reaction: $\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{ZnCl}_2\text{(aq)}$
Which set of reaction conditions produces $\text{H}_2\text{(g)}$ at the fastest rate?
- a 1.0-g lump of Zn(s) in 50 mL of 0.5 M HCl(aq) at 20 degrees Celsius
 - a 1.0-g lump of Zn(s) in 50 mL of 0.5 M HCl(aq) at 30 degrees Celsius
 - 1.0 g of powdered Zn(s) in 50 mL of 1.0 M HCl(aq) at 20 degrees Celsius
 - 1.0 g of powdered Zn(s) in 50 mL of 1.0 M HCl(aq) at 30 degrees Celsius
- 2) At 20°C, a 1.2-gram sample of Mg ribbon reacts rapidly with 10.0 milliliters of 1.0 M HCl(aq) .
Which change in conditions would have caused the reaction to proceed more **slowly**?
- increasing the initial temperature to 25°C
 - decreasing the concentration of HCl to 0.1 M
 - using 1.2 g of powdered Mg
 - using 2.4 g of Mg ribbon

Writing Chemical Equations: Representing a physical or chemical change

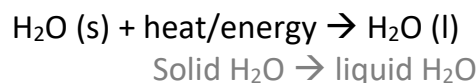
For chemical changes, there are two terms you need to be familiar with:

- **Reactants**= the substances that you _____ in a chemical reaction
- **Products**= the substances that you _____ in a chemical reaction

Example: Photosynthesis



Example: melting of ice



*Note: don't use the word "reactant" or "product" here as no new substance was formed.

Important Symbols in Chemical Equations:

Symbol	Meaning
	Separates 2 or more reactants or products (“and” or “combine” or “reacts”)
	Separates reactants from products (“yield/produce”)
	Refers to a substance in a solid state
	Refers to a substance in a liquid state
	Refers to a substance in gas state
	Refers to a substance that is dissolved in water (aqueous solution)

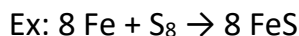
Practice Writing Chemical Equations: Don't forget to include energy/heat is a reactant/product and the phases of each substance! Also, don't forget about your diatomic elements!

- 1) When sodium metal reacts with iron (III) chloride (FeCl_3), iron metal and sodium chloride (NaCl) are formed.
- 2) The Haber Process is an exothermic reaction in which nitrogen (N_2) and hydrogen (H_2) gas are combined to produce ammonia (NH_3) gas
- 3) When dissolved barium chloride (BaCl_2) reacts with dissolved potassium sulfate (K_2SO_4) in water, barium sulfate (BaSO_4) precipitates and aqueous potassium chloride (KCl) is made.
- 4) When sucrose ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$) burns in oxygen gas (O_2), carbon dioxide, water and heat are produced.
- 5) Hydrogen peroxide (H_2O_2) breaks down to form water and oxygen gas (O_2).
- 6) When dissolved calcium hydroxide ($\text{Ca}(\text{OH})_2$) reacts with sulfuric acid (H_2SO_4), a calcium sulfate (CaSO_4) precipitate is formed along with water and heat
- 7) When dry ice (CO_2) is placed on a countertop at room temperature, it sublimates.

Types of Reactions:

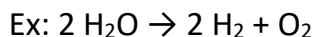
Synthesis: 2 or more elements combine to make a compound;

General form: _____



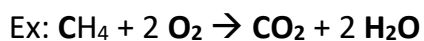
Decomposition: A compound breaks down into 2 or more simpler substances;

General Form: _____



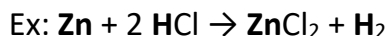
Combustion: A compound containing carbon and hydrogen is burned in the presence of oxygen, and the products are carbon dioxide & water vapor

General Form: _____



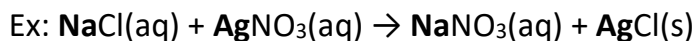
Single Replacement: 1 element replaces another element in a compound;

General Form: _____



Double Replacement: 2 elements of 2 different compounds switch places with one another;

General Form: _____



Neutralization: A double replacement reaction that takes place when an acid and a base react to form a metal salt and water. (This is also known as an Acid-Base Reaction)

General Form: _____



Type of Reaction	Chemical Equation
	$2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$
	$2 \text{ C}_6\text{H}_{14} + 19 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 14 \text{ H}_2\text{O}$
	$2 \text{ K} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ KOH} + \text{H}_2$
	$4 \text{ Cr} + 3 \text{ O}_2 \rightarrow 2 \text{ Cr}_2\text{O}_3$
	$\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2 + \text{Na}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2 \text{ NaC}_2\text{H}_3\text{O}_2$
	$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

Homework

- 1) Decide whether the following are physical changes (PC) or chemical changes (CC)
- | | |
|---|--|
| a) Salt dissolves in water _____ | f) Ethyl alcohol evaporates _____ |
| b) Hydrochloric acid reacts with magnesium to produce hydrogen gas ____ | g) Paper burns _____ |
| c) A piece of copper is cut in half _____ | h) Fireworks explode _____ |
| d) A sugar cube is ground up _____ | i) Sodium and potassium react violently with water _____ |
| e) Iron rusts _____ | j) Water freezes _____ |

- 2) Which of the following phase changes are endothermic?
- | | |
|----------------------------|---------------------------------|
| A. Melting and freezing | C. Condensation and sublimation |
| B. Melting and evaporating | D. Condensation and deposition |

- 3) A 5.0-gram sample of zinc and a 50-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
- | | |
|-----------------------------------|----------------------------------|
| A. a zinc strip and 1.0 M HCl(aq) | C. zinc powder and 1.0 M HCl(aq) |
| B. a zinc strip and 3.0 M HCl(aq) | D. zinc powder and 3.0 M HCl(aq) |

**Name another factor that could increase the rate of this reaction: _____

- 4) **Write out the chemical equation & identify the reaction type** for the following reactions.

- a) When a solution of BeCl_2 reacts with an AgNO_3 solution, aqueous $\text{Be}(\text{NO}_3)_2$ and AgCl powder are made.
- b) When isopropanol ($\text{C}_3\text{H}_8\text{O}$) burns in oxygen gas (O_2), carbon dioxide, water, and heat are produced.
- c) When dissolved NaOH reacts with H_2SO_4 , an exothermic reaction occurs in which aqueous sodium sulfate (Na_2SO_4) and water are formed.
- d) When fluorine gas (F_2) is put into contact with calcium metal, CaF_2 powder is created in an exothermic reaction.
- e) When sodium metal reacts with FeCl_2 , iron metal and sodium chloride are formed.